

**PATENT SPECIFICATION**

**TITLE: IMPROVED METHODS AND PROCESSES OF HYDROGEN PEROXIDE PRODUCTION**

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**RELATED APPLICATION DATA**

This application claims priority based on a provisional application, U.S. 60/390,976.

**Field of the Invention**

The present invention relates to simplified processes for the preparation of hydrogen peroxide ( $\text{H}_2\text{O}_2$ ).  $\text{H}_2\text{O}_2$  is a known oxidizer and disinfectant that is used in many industrial processes having many uses in the pharmaceutical, electronic, food and water purification industries.

The present invention presents the use of sulfuric acid ( $\text{H}_2\text{SO}_4$ ) as a catalyst utilizing water and electricity as the only raw materials for the production of  $\text{H}_2\text{O}_2$ .  $\text{H}_2\text{O}_2$  is an ideal oxidizer and disinfectant in water purification systems, especially drinking water purification. All other disinfectants create disinfection by-products upon their reaction with Natural Organic Matter (NOM) in the water. Many disinfection by-products currently produced from chlorine, chlorine dioxide and ozone are known toxins, carcinogens and teratogens in drinking water. In contrast, other than the oxidized organic molecule,  $\text{H}_2\text{O}_2$  after reaction breaks down to water ( $\text{H}_2\text{O}$ ) and Oxygen ( $\text{O}_2$ ).

Current  $\text{H}_2\text{O}_2$  manufacturing processes utilize the Anthraquinone Process (AP) to produce hydrogen peroxide from 2-ethyl anthraquinone. This is a two stage process; wherein the first stage 2-ethyl anthraquinone reacts with Hydrogen ( $\text{H}_2$ ) over a lead catalyst to produce 2-ethyl dihydroanthraquinone and in the second stage, the 2-ethyl dihydroanthraquinone from the first stage, reacts with  $\text{O}_2$  to produce  $\text{H}_2\text{O}_2$  and 2-ethyl anthraquinone for recycle. This process requires the handling of two explosive chemicals,  $\text{H}_2$  and  $\text{O}_2$ , as well as the handling of hazardous anthraquinone chemistries. AP requires an industrial manufacturing environment for the production of  $\text{H}_2\text{O}_2$  due to the safety issues associated with: the chemicals, the manufacturing process and the handling/transportation of  $\text{H}_2\text{O}_2$ .

### Background of the Invention and Description of the Prior Art

In the nineteenth century,  $\text{H}_2\text{O}_2$  was produced exclusively on a commercial scale by the reaction of an acid upon barium peroxide, which is readily formed by heating barium oxide in air. The barium peroxide was then usually reacted with sulfuric acid to produce  $\text{H}_2\text{O}_2$  along with a solid precipitate of barium sulfate.  $\text{H}_2\text{O}_2$  is normally produced in the laboratory by the reaction of sodium peroxide with water or dilute hydrochloric acid. In this reaction, sodium oxide reacts with either  $\text{H}_2\text{O}$  or hydrochloric acid to form  $\text{H}_2\text{O}_2$  along with either sodium hydroxide in the case of  $\text{H}_2\text{O}$  or sodium chloride in the case of hydrochloric acid.

Until the mid 20'th century,  $\text{H}_2\text{O}_2$  was produced by the electrolysis of  $\text{H}_2\text{SO}_4$ , the Sulfuric Acid Process (SAP). SAP is a two stage process wherein the first stage  $\text{H}_2\text{SO}_4$  is electrochemically converted to  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$ . In the second stage, the  $\text{H}_2\text{S}_2\text{O}_8$  from the first stage is reacted with  $\text{H}_2\text{O}$  to form  $\text{H}_2\text{O}_2$  and  $\text{H}_2\text{SO}_4$  for recycle. At that time, this process was not as economical or as safe to operate as AP. At that time, SAP required the energy of electrolysis

along with two stages of distillation, wherein the first stage of distillation required the separation of an explosive chemical,  $H_2$ , and the second stage of distillation required the separation of  $H_2O_2$  from very corrosive  $H_2SO_4$ . AP replaced SAP due to both economical and safety issues.

However, AP presents other issues, the first of which is safety.  $H_2O_2$  is a very hazardous chemical to store and to transport. Once produced,  $H_2O_2$  requires specialized equipment in storage and in transportation.  $H_2O_2$  at concentrations of over 20% can react violently with organic materials and with acids.  $H_2O_2$  in concentrations of over 50% can explosively decompose or react with any electron acceptor material, such as acids, or organic materials. Another issue in relation to the AP production of  $H_2O_2$  is purity. The AP process is rather complex. The steps in this process to produce high grade  $H_2O_2$  include: hydrogenation, oxidation, extraction, solvent scrubbing, settling, cation resin treatment, carbon treatment, cation resin treatment, distillation, absorption and dilution. The product after the initial cation resin treatment is very pure except for the presence of organic contaminants, which are one of the key quality parameters for electronics grade peroxide, namely, total carbon.

Perhaps the single problem most experienced in the electronics industry is contamination during the manufacture of microcircuit devices, such as chips and wafers. As these devices become more complex and smaller in dimension, sensitivity to contamination becomes more acute. Contamination is a problem because contaminants, in the form of solid particles, can open or short a circuit, affect photolithographic reproduction, alter electrical properties and even damage the crystal structure of modern electronic devices.

In addition to the electronics industry, there are many uses for  $H_2O_2$  in the water purification industry. Oxidation and disinfection processes with  $H_2O_2$  produce much purer water than that with other oxidizers. Chlorine, chlorine dioxide, ozone, chloramines and bromine all

produce by-products of disinfection that are either: toxic, carcinogenic or teratogenic. As such, these by-products are health issues to humanity, plant and animal life. In contrast, the products of oxidation with  $H_2O_2$  are usually limited to oxygen ( $O_2$ ), carbon dioxide ( $CO_2$ ) and  $H_2O$ , along with the organic oxidation product. However, most water purification facilities are not interested  
5 in the storage and handling of  $H_2O_2$  due to the safety issues associated with  $H_2O_2$ .

This instant invention proposes end-use on-site production of  $H_2O_2$  utilizing a newly found SAP process instead of AP. This instant invention presents an improvement upon historic SAP by performing separation with membranes instead of by distillation. The utilization of membranes, either inorganic or organic or a combination thereof, would provide the ability of  
10 on-site production of  $H_2O_2$ , thereby eliminating the need for AP, as well as the transportation, storage and handling of  $H_2O_2$ .

Membrane technologies have been available since the early 1990's. Early versions of this technology were of organic form only. As of late, inorganic materials as known in the art have been incorporated in many applications. As of late, membrane technologies have been  
15 improved to incorporate electrolysis directly onto the membrane. This is accomplished by various methods known in the art, which make the membrane conductive; this conductivity can be accomplished in both inorganic and organic membranes.

In fuel cells, U.S. Pat. No. 4,490,445 teaches a solid oxide electrochemical energy converter comprising alternating layers of solid oxide electrolyte plates and electrical conductor  
20 plates. Each electrolyte plate includes a coating of a porous oxidizer electrode on a first surface of the electrolyte and a coating of a porous fuel electrode on a second surface of the electrolyte. Each conductor plate includes grooved networks formed by ridges which define gas passages on both surfaces of the conductor plate, such ridges being in electrical contact with electrode

coatings on next adjacent electrolytes. Each conductor plate also possesses a means for tapping electricity from or introducing electricity into the converter. The conductor plates also possess circumferential ridges arranged along the edges of the conductor plate to define gas seals, the ridges being in contact with surface coatings on next adjacent electrolyte plates in which surface coatings possess the same composition as that of the electrode coatings.

U.S. Patent No. 4,791,079 discloses two-layer conducting catalytic ceramic membranes which are suitable for use in a variety of hydrocarbon conversion reactions. The two-layer membrane possesses one layer formed of an impervious mixed ion and electronic conducting ceramic such as yttria stabilized zirconia which is doped with sufficient  $\text{CeO}_2$  or  $\text{TiO}_2$  to impart electron conducting characteristics to the ceramic. A second layer associated with mixed conducting impervious ceramic is a porous ion conducting layer containing a selective hydrocarbon oxidation catalyst.

A variety of methods have been presented for the separation of oxygen from air. A solid electrolyte oxygen pump formed from a plurality of solid-state membranes is presented in U.S. Patent No. 4,877,506. The oxygen pump possesses electrodes which are shaped to form a plurality of linear, parallel channels on facing surfaces of the electrolyte. The air feed is introduced into the channels formed of the air electrode. Oxygen formed during operation of the device is removed by passage through the electrolyte via channels formed of the oxygen electrode or anode. A monolithic array is formed by situating an interconnecting material between adjacent cells to form a stack of cells. U.S. Patent No. 5,034,023 discloses ceramic honeycomb structures which are capable of separating oxygen from an oxygen-containing gaseous mixture. The channeled honeycombs are formed from a solid electrolyte having at least some of the honeycomb channels sealed at one of its faces. The oxygen-containing gas is

introduced into a first set of channels at one face of the honeycomb, a first voltage is applied to the interior walls of the channels and a second voltage is applied to the interior walls of the second set of remaining channels thereby creating an electrical potential across the ceramic material separating adjacent channels of the two sets. The electrical potential drives oxygen ions through the channel walls releasing molecular oxygen into the second set of channels which can be collected. U.S. Patent No. 5,045,169 discloses an electrochemical device capable of generating oxygen from air upon the application of an electrical current, where a plurality of adjacent electrochemical cells are electrically connected in series, each cell containing an inner, porous oxygen electrode; a dense, solid oxide electrolyte capable of transporting oxygen ions partly disposed on top of the inner electrode and partly disposed between inner electrodes of adjacent cells; an outer porous air electrode disposed on top of the electrolyte; and separate, dense, electronically conductive segments of interconnection material disposed between adjacent cells, the interconnection electrically and physically connecting the outer air electrode from one cell to the inner oxygen electrode from an adjacent cell, the device having gas impermeable, dense, contacting segments of electrolyte and interconnection material between inner electrode of adjacent cells. U.S. Patent No. 5,240,480 discloses representative solid-state membranes for separating oxygen from oxygen-containing gaseous mixtures. These membranes comprise a multi-component metallic oxide porous layer having an average pore radius of less than about 10 micrometers and a multi-component metallic oxide dense layer having no connected through porosity wherein the porous layers and dense layers are contiguous and such layers conduct electrons and oxygen ions at operating temperatures.

U.S. Patent No. 5,356,728 and European Patent Application WO 94/24065 disclose cross-flow electrochemical reactor cells formed from multicomponent metallic oxides of the

perovskite structure which demonstrate electron conductivity and oxygen ion conductivity at elevated temperatures. Such cells are useful in carrying out partial oxidation reactions of organic compounds to form added-value products and separating oxygen from oxygen-containing gaseous mixtures. The cross-flow reactor cells of U.S. Patent No. 5,356,728 comprise either a hollow ceramic blade positioned across a gas stream flow containing one or more channels for flow of gas streams or a stack of crossed hollow ceramic blades containing one or more channels for flow of gas streams. Each channel has at least one channel wall disposed between a channel and a portion of an outer surface of the ceramic blade or a common wall with adjacent blades in a stack comprising a gas impervious multi-component metallic oxide, typically of a perovskite structure, which exhibits electron conductivity and oxygen ion conductivity at elevated temperatures. Thus, the channels are contiguous to the outer surface of the ceramic blade which is formed from the multi-component metallic oxide.

None of these patents teach or suggest electrolysis and membrane separation or electrolytic membrane separation as methods of  $H_2O_2$  preparation.

In  $H_2O_2$  manufacturing, membranes have been discussed as methods of  $H_2O_2$  purification. U.S. Patent Nos. 4,879,043 and 6,333,018 present the use of reverse osmosis membrane technology as a final purification step in the production of  $H_2O_2$  manufactured by AP. U.S. Patent Nos. 5,215,665; 5,262,058 and 5,906,738 present the use of reverse osmosis membrane technology in combination with cationic resin technology as final purification steps in the production of  $H_2O_2$  manufactured by AP. U.S. Patent Nos. 5,851,042 and 6,113,798 present the use of converting contaminant particles by reacting said particles with micro-ligands, then separating said reaction products with membranes as a final purification step in the production of

H<sub>2</sub>O<sub>2</sub> manufactured by AP. None of these patents teach or suggest electrolysis and membrane separation or electrolytic membrane separation as methods of H<sub>2</sub>O<sub>2</sub> preparation.

U.S. Patent No. 5,800,796 presents an electrochemical reactor wherein O<sub>2</sub> and H<sub>2</sub> are reacted across a conductive membrane containing reducing catalysts forming H<sub>2</sub>O<sub>2</sub>. This novel process eliminates AP while simplifying the process of H<sub>2</sub>O<sub>2</sub> production. However, the potential for contamination of H<sub>2</sub>O<sub>2</sub> with heavy metals from the reducing catalyst is significant. Heavy metals contamination eliminates the potential use of H<sub>2</sub>O<sub>2</sub> in either the production of micro-circuitry or water purification. In addition, the potential safety issues from the reaction of very explosive O<sub>2</sub> and H<sub>2</sub> in an electrolytic environment preclude the potential use of this process at the end-use site.

All of these applications and many other industry applications of organic and inorganic membranes have shown the potentiality of membrane technologies. However, none of these applications or any other industrial application has proposed the use of membranes to improve the manufacturing process or the safety associated with and/or the handling of H<sub>2</sub>O<sub>2</sub>.

### Summary of the Invention

A primary object of the invention is to devise an effective, efficient and economically feasible process for producing H<sub>2</sub>O<sub>2</sub>.

Another object of the invention is to devise an effective, efficient and economically feasible process for producing H<sub>2</sub>O<sub>2</sub>, wherein the safety in handling of H<sub>2</sub>O<sub>2</sub> is improved.

Another object of the invention is to devise an effective, efficient and economically feasible process for producing H<sub>2</sub>O<sub>2</sub>, wherein transportation of H<sub>2</sub>O<sub>2</sub> is not required.

Another object of the invention is to devise an effective, efficient and economically feasible process for producing  $\text{H}_2\text{O}_2$ , wherein significant storage of  $\text{H}_2\text{O}_2$  is not required.

Another object of the invention is to devise an effective, efficient and economically feasible process for producing  $\text{H}_2\text{O}_2$ , wherein high purity  $\text{H}_2\text{O}_2$  is manufactured.

5 Another object of the invention is to devise an effective, efficient and economically feasible process for producing  $\text{H}_2\text{O}_2$ , utilizing  $\text{H}_2\text{SO}_4$  as a catalyst along with water and electricity as the raw materials.

Another object of the invention is to devise an effective, efficient and economically feasible process for handling  $\text{H}_2\text{O}_2$  at water treatment plants so that  $\text{H}_2\text{O}_2$  is an attractive  
10 alternate to other oxidizers and disinfectants utilized at water treatment plants.

Additional objects and advantages of the invention will be set forth in part in a detailed description which follows, and in part will be obvious from the description, or may be learned by practice of the invention.

## 15 **Brief Description of the Drawings**

Figure 1 illustrates in block diagram form a general description of the proposed methods and processes of the instant invention wherein,  $\text{H}_2\text{SO}_4$  is converted by electrolysis to  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$  and wherein, said  $\text{H}_2\text{S}_2\text{O}_8$  is reacted with  $\text{H}_2\text{O}$  to form  $\text{H}_2\text{O}_2$  and  $\text{H}_2\text{SO}_4$  and wherein, said  $\text{H}_2\text{SO}_4$  from  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2\text{O}$  is recycled for electrolysis into  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$ .

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## **Description of the Preferred Embodiments**

The instant invention: improves the purity of  $\text{H}_2\text{O}_2$ , simplifies the manufacture of  $\text{H}_2\text{O}_2$ , eliminates the need to store large volumes of concentrated  $\text{H}_2\text{O}_2$  and eliminates the need to

transport  $\text{H}_2\text{O}_2$ . The instant invention presents the use of sulfuric acid ( $\text{H}_2\text{SO}_4$ ) as a catalyst utilizing water ( $\text{H}_2\text{O}$ ) and electricity as the only raw materials for the production of  $\text{H}_2\text{O}_2$ . The instant invention utilizes membrane technology in combination with electrolysis to produce  $\text{H}_2\text{O}_2$  at the site of the end-user, whereupon  $\text{H}_2\text{O}_2$  can be utilized with minimal storage and no transportation. The process of the instant invention, the New Sulfuric Acid Process (NESAP), is presented as a two stage process, wherein the first state  $\text{H}_2\text{SO}_4$  is most preferably electrochemically converted to  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$ . The  $\text{H}_2$  is separated and preferably used in a fuel cell to generate electricity. Said electricity for electrolysis is preferably at least partially used in the generation of  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$  from  $\text{H}_2\text{SO}_4$ . In the second stage, the  $\text{H}_2\text{S}_2\text{O}_8$  from the first stage is reacted with  $\text{H}_2\text{O}$  to form  $\text{H}_2\text{O}_2$  and  $\text{H}_2\text{SO}_4$ .  $\text{H}_2\text{SO}_4$  is preferably recycled to the first stage for electrochemical conversion, again, to  $\text{H}_2\text{S}_2\text{O}_8$  and  $\text{H}_2$ . Since it can be rather cost ineffective to perform separations which are precise, separation in this invention is to be defined to mean substantial separation or separation of such percentage that the chemical process and the application of  $\text{H}_2\text{O}_2$  is as intended. It is to be understood that within a stage of separation, whether that separation be of distillation or of membrane technology or a combination thereof, that improved separation can be accomplished with multiple stages of separation as compared to a single stage.

It is preferred that said electrolysis in the first stage be performed with at least one of a conductive organic membrane and a conductive inorganic membrane. It is most preferred that said organic and/or said inorganic membrane be conductive as is known in the art. It is preferred to perform electrolysis in the first stage and that the separation of  $\text{H}_2$  from  $\text{H}_2\text{S}_2\text{O}_8$  and unreacted  $\text{H}_2\text{SO}_4$  is performed by at least one of membrane separation and distillation. It is a preferred embodiment to perform electrolysis with an electrode in the first stage. It is preferred that said

electrode be made of a corrosion resistant material. It is preferred that said electrode material be made at least one of: zirconium, hastelloy, titanium, ceramic and any combination thereof. It is most preferred that the separation of  $H_2$  from  $H_2S_2O_8$  and unreacted  $H_2SO_4$  be performed with a membrane. It is preferred that the separation of  $H_2$  from  $H_2S_2O_8$  and unreacted  $H_2SO_4$  be performed with distillation. It is preferred that the separation of  $H_2S_2O_8$  from unreacted  $H_2SO_4$  be performed with at least one of distillation and membrane separation. It is preferred that the separation of  $H_2$  from  $H_2S_2O_8$  and unreacted  $H_2SO_4$  be performed with at least one of distillation and membrane separation.

It is most preferred that, in the second stage, the separation of at least one of  $H_2O_2$  and unreacted  $H_2O$  from at least one of  $H_2SO_4$  and unreacted  $H_2S_2O_8$  be performed with a membrane. It is most preferred that the separation of  $H_2SO_4$  from unreacted  $H_2S_2O_8$  in the second stage be performed with a membrane. It is an embodiment that the separation of at least one of  $H_2O_2$  and unreacted  $H_2O$  in the second stage from at least one of  $H_2SO_4$  and unreacted  $H_2S_2O_8$  be performed with distillation. It is an embodiment that the separation of  $H_2SO_4$  from  $H_2S_2O_8$  in the second stage be performed with distillation. It is most preferred that the second stage be separated into two portions, wherein at least one of  $H_2O_2$  and  $H_2O$  are separated from at least one of  $H_2SO_4$  and unreacted  $H_2S_2O_8$  in the first portion and  $H_2SO_4$  is separated from unreacted  $H_2S_2O_8$  in the second portion. It is most preferred that said separated  $H_2SO_4$  in the second stage flow to the first stage.

It is most preferred that an excess amount of  $H_2O$  be added to the second stage so that the produced  $H_2O_2$  is dilute upon formation to a concentration which maximizes safety in a given end-use application. It is preferred that the  $H_2O_2$  in the second stage be diluted with  $H_2O$  upon

separation of said  $\text{H}_2\text{O}_2$  from at least one of  $\text{H}_2\text{SO}_4$  and/or  $\text{H}_2\text{S}_2\text{O}_8$  to a concentration which maximizes safety in a given end-use application.

### Physical Properties

Water,  $\text{H}_2\text{O}$ , is clear with no color having a freezing point of  $0^\circ\text{C}$  and a boiling point of  $100^\circ\text{C}$ .

$\text{H}_2\text{SO}_4$  has a freezing point of  $10^\circ\text{C}$  and a boiling point of  $338^\circ\text{C}$ .

$\text{H}_2\text{O}_2$  has a freezing point of  $0^\circ\text{C}$  and a boiling point of  $150^\circ\text{C}$ .

While  $\text{H}_2$  is a colorless gas at ambient conditions,  $\text{H}_2\text{S}_2\text{O}_8$  is a yellowish to brown solid below  $65^\circ\text{C}$  and decomposes above its melting point; however,  $\text{H}_2\text{SO}_4$  is a good solvent for  $\text{H}_2\text{S}_2\text{O}_8$ . In the instant invention it is preferred that stage 1, the transfer line from stage 1 to stage 2 and stage 2 be insulated having a temperature control loop to maintain operating temperatures. It is preferred that the operating temperature of stage 1 and the transfer line from stage 1 to the  $\text{H}_2\text{S}_2\text{O}_8/\text{H}_2\text{O}$  reaction chamber be approximately between  $20$  and  $70^\circ\text{C}$ , and most preferably between  $50$  and  $65^\circ\text{C}$ . Further, it is preferred that the electrolysis in stage 1 not carry to completion so that residual  $\text{H}_2\text{SO}_4$  can be used as a solvent for  $\text{H}_2\text{S}_2\text{O}_8$  transferred to Stage 2. Recycle of  $\text{H}_2\text{O}$  from stage 2 to stage 1 in the  $\text{H}_2\text{SO}_4$  recycle line will result in the electrolysis of  $\text{H}_2\text{O}$  in stage 1, thereby reducing system efficiency while creating  $\text{O}_2$  in the system or sending  $\text{O}_2$  to the  $\text{H}_2$  side of the fuel cell. In light of this operating possibility, in the instant invention it is preferred that stage 2 have an operating temperature of between  $100$  to  $170^\circ\text{C}$  and most preferably between  $105$  to  $115^\circ\text{C}$ , along with an  $\text{O}_2$  purge in the  $\text{H}_2\text{S}_2\text{O}_8/\text{H}_2\text{SO}_4$  line from stage 1 to stage 2. It is preferred to send the  $\text{O}_2$  purge to the air (oxidizer) side of the fuel cell.

Said temperature control loops can be any heating/cooling loop as is known in the art. It is preferred to heat stage 1 and/or stage 2 with an electrical resistant circuit and to cool stage 1

and/or stage 2 with a  $\text{H}_2\text{O}$  jacket. It is most preferred to use the heat of electrolysis to heat stage 1 and the heat of reaction of  $\text{H}_2\text{O}$  with  $\text{H}_2\text{S}_2\text{O}_8$  to maintain operating temperature in stage 2 and/or stage 1. It is an embodiment to heat stage 1 and/or stage 2 with a steam jacket. It is most preferred to manage the temperature in stage 2 by the temperature of the water/steam added to the reaction chamber for the  $\text{H}_2\text{S}_2\text{O}_8/\text{H}_2\text{O}$  reaction to  $\text{H}_2\text{O}_2$ . Stage 1 operating temperature is most preferably maintained by the cooling of  $\text{H}_2\text{SO}_4$  from stage 2 to stage 1. Said cooling of  $\text{H}_2\text{SO}_4$  is most preferably accomplished with ambient convective heat transfer. Cooling of  $\text{H}_2\text{SO}_4$  from stage 2 to stage 1 is preferably accomplished with a water jacket on the  $\text{H}_2\text{SO}_4$  transfer line from stage 2 to stage 1. Once operating temperatures are obtained, cooling will often be required to remove the heat energy of electrolysis generated in stage 1.

It is preferred that the production of  $\text{H}_2\text{O}_2$  by this instant invention be performed at the end-use location to minimize or eliminate at least one of storage and/or vehicular transportation of  $\text{H}_2\text{O}_2$ . It is most preferred to utilize  $\text{H}_2\text{O}_2$  manufactured by this instant invention in the manufacture of electric circuit devices. It is most preferred to utilize  $\text{H}_2\text{O}_2$  manufactured by this instant invention in the purification of water.

### Example 1

Concentrated sulfuric acid is placed in a flask having three openings, one on each side and one in the middle. Inside one side opening is placed a rod of titanium; this opening is sealed with a rubber stopper. The sulfuric acid solution is heated to  $50\text{ }^\circ\text{C}$ , as measured by a thermometer in the middle opening sealed with a rubber stopper. Inside the other side opening is placed a rod of titanium; this opening is left open. A 12 volt dc car battery is connected to the electrodes, with the cathode at the open port and the anode at the closed port. Within about a

minute the concentrated sulfuric acid solution begins to turn yellow. The open port of the flask is lit with a match. The open port obviously contains hydrogen. Reaction is carried out until a full yellowish color is obtained in the flask. The electrodes are removed.

## 5     **Example 2**

The resultant solution from Example 1 is then slowly reacted with water until the solution becomes clear again. Once clear, the solution is heated to boiling, which occurs at about 100 °C and increases in intensity at about 150 °C. The distilled vapors are obviously a combination of water and hydrogen peroxide. The remaining liquid has a pH of less than 1.0, being sulfuric  
10     acid.

## **Example 3**

Example 2 is repeated. This time two flask openings are sealed, each sealed with a glass stopper. A Teflon tube is placed with one end on the third flask opening and the other end of the hose in a  
15     beaker of water. The flask is heated to 155 °C and boiled until boiling stops. The resultant distillate/water mixture is then poured on a pair of old leather shoes. An exothermic reaction takes place wherein, the shoes begin to smoke. The exothermic reaction on leather proves the resultant aqueous solution to contain hydrogen peroxide.

20     Certain objects are set forth above and made apparent from the foregoing description, tables, drawings and examples. However, since certain changes may be made in the above description, tables, drawings and examples without departing from the scope of the invention, it is intended that all matters contained in the foregoing description, tables, drawings and examples

shall be interpreted as illustrative only of the principles of the invention and not in a limiting sense. With respect to the above description, tables, drawings and examples then, it is to be realized that any descriptions, tables, drawings and examples deemed readily apparent and obvious to one skilled in the art and all equivalent relationships to those stated in the tables,  
5 drawing and examples and described in the specification are intended to be encompassed by the present invention.

Further, since numerous modifications and changes will readily occur to those skilled in the art, it is not desired to limit the invention to the exact construction and operation shown and described, and accordingly, all suitable modifications and equivalents may be resorted to, falling  
10 within the scope of the invention. It is also to be understood that the following claims are intended to cover all of the generic and specific features of the invention herein described, and all statements of the scope of the invention which, as a matter of language, might be said to fall in between.

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